



UNITED NATIONS SCHOOL I.E.D.
PEI: COMPREHENSIVE TRAINING OF COMPETENT ENTREPRENEURIAL LEADERS, WITH
DEMOCRATIC, TECHNOLOGICAL, CULTURAL AND SPORTS PRINCIPLES
MOTTO: "EDUCATION, SCIENCE, CULTURE AND SPORT TO TRANSCEND"

PREPARATION WORKSHOP FOR THE FIRST PERIOD
CHEMISTRY
SEVENTH GRADE
TEACHER HEISEL QUESADA

The preparation workshop must be carried out in the Chemistry notebook as a requirement to take the competency test

Delivery date: April 1

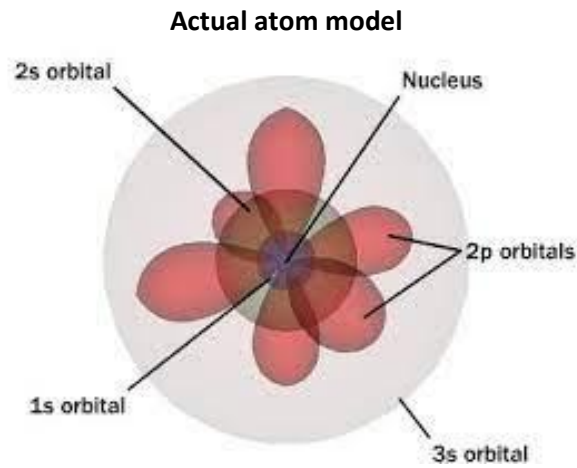
The mass number (A) of an atom is the number of protons (p+) positive particles and neutrons (n°) neutral or uncharged particles found in the atomic nucleus. Z represents atomic number or number of protons in a neutral atom equal to the number of electrons or negative particles which are in constant motion

$$A = Z + n$$

The table below shows some characteristics for three atoms, P, Q, R

	P	Q	R
Z	19		25
A		20	52
N	18	10	

1. Define the following terms
 - a. Masic number
 - b. Protons
 - c. Neutrons
 - d. Electrons
2. Fill in the box with the missing data
3. Describes the steps to calculate the number of neutrons (n) in an atom
4. Design the electron distribution for P, Q, R atoms
5. What if the atomic nucleus had no neutrons?
6. Propose a scheme for the atoms P, Q, R



7. Write about the advances that the current atomic model has made compared to the models of Dalton, Thomson, Rutherford, and Bohr
8. What would be the atom represented in the diagram? Justify your answer
9. Is it possible that there are more than 118 chemical elements? Justify your answer
10. Describes how the periodic table would be represented by the current atomic model