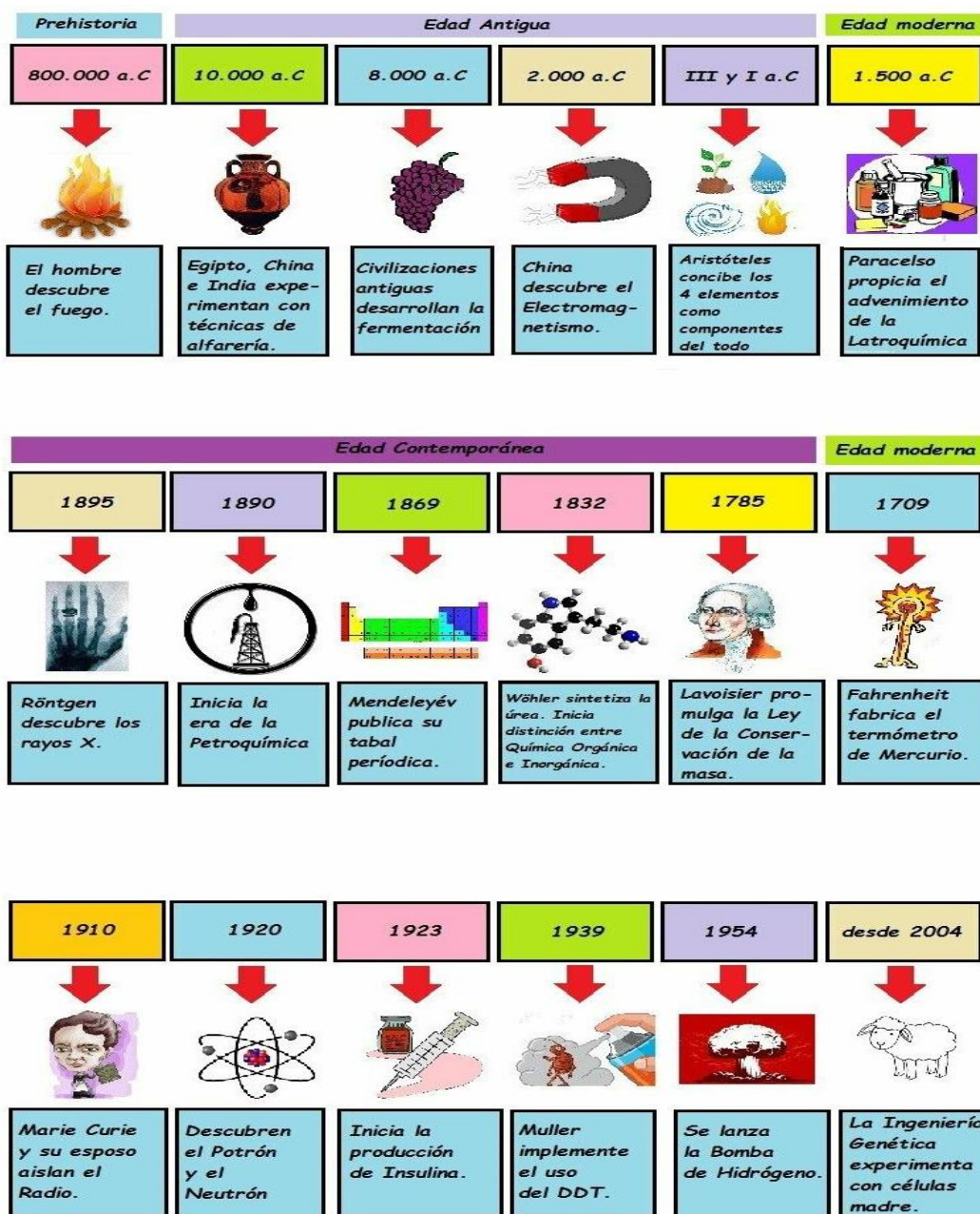




PREPARATION WORKSHOP
THIRD PERIOD
CHEMISTRY
TENTH GRADE
TEACHER HEISEL QUESADA

Delivery date November 3, 2023



From the timeline answer:

1. What are the main events that occurred in
 - a. Old age
 - b. Contemporary age
 - c. Modern age
2. What do you think is the most important discovery in the history of chemistry? Justify your answer
3. Make a concept map with the history of Chemistry
4. Draw the following atoms with nucleus, levels and orbitals
 - a. 65
 - b. 32
 - c. 9
 - d. 93

5. Draw with different colors the periodic table, indicating the groups, levels, zones, no metals, noble gases and rare earth

5. Draw the periodic table and locate

a. Z: 49, 30, 20

b. Germanio, Uranio, Francio

6. For the previous elements calculate all the atomic characteristics

7. Complete the following information

element	simbol	Atomic number (Z)	group	Electro negativity	Valence e-	valences	Oxidation numbers
sodium							
	Chlorine						
		81					
Galium							
	Hg						

8. Realice the concep map with the subjet "Chemical bond"

9. Graphically and numerically calculate the type of bond formed between

- a. Sr - N
- b. Cr - As
- c. Re - Cl

10. Organize the following electronegativities from lowest to highest, writing the electronegativity of each one

Br, At, Cl, U, Fe, Ni, Na, Ra, I, N, Po, Mn, Bi, Pb, Th, Be, O, S

11. Write between which pairs of elements ionic and covalent bonds would form. Realice the procedure

12. Calculate the chemical percentages in a solution with 200g of solute and 6L of solution

13. Calculate the solute of 3500 ml CO₂, percentage 25%

14. A solution has 38% of H₂SO₃. Calculate the total solution with 430 g of solvent

Complete the chart (with procedures)

	sustance	Molecular weight	grams	moles	molécules
15	U(OH) ₃				
16	Th(CLO ₂) ₂				
17	MnBr ₇				
18	HAsO ₂				
19	Cr ₂ (SO ₂) ₃				
20	Au(OH) ₂				

Complete the chart (with procedures)

	Sustance	Molecular weight	Solute(mol)	Solution (L)	Molarity (mol/L)
21	Re(OH) ₇		400 g		
22	Cl ₂ O ₇			4600ml	
23	HBrO				0.6
24	Ra(ClO) ₂		9000 molec		
25	FeO			6 L	

Complete the chart (with procedures)

	Sustance	Molecular weight	Equivalent weight	Solute (eq-g)	Solution (L)	Normality (Eq-g/L)
26	CO ₂					8
27	HIO ₃				3600 ml	
28	Ca(NO ₃) ₂			45		
29	SnO				7 L	
30	Au(OH) ₃					0.003